•	The reaction of methane and water is one way to prepare hydrogen for use as a fuel. $CH_4(g) \ + \ H_2O(g) \ \to \ CO(g) \ + \ 3H_2(g)$	Marks 3
	Which compound is the limiting reactant if you begin with 995 g of methane and 2510 g of water?	
	Answer:	_
	What mass of the excess reactant remains when the reaction is completed?	
	Answer:	