Marks
4

- An unknown compound contains carbon and hydrogen only. If 0.0956 g of the compound is burned in oxygen, 0.300 g of CO_2 and 0.123 g of H_2O are isolated. What is the unknown compound's empirical formula?

Answer:

If its molar mass is found to be 70.1 g mol^{-1}, what is its molecular formula?

Answer:

1

- What amount (in mol) of chloride ion is contained in 100 mL of 0.25 M magnesium chloride solution?

Answer:

1

- If 25.0 mL of 1.50 M hydrochloric acid is diluted to 500 mL, what is the molar concentration of the diluted acid?

Answer: