**Marks
6**

- Lead(II) iodide precipitates when 0.080 M lead(II) nitrate solution (150.0 mL) is added to 0.080 M potassium iodide solution (50.0 mL). Write a balanced ionic equation for the reaction that occurs.

What amount (in mol) of lead(II) iodide precipitates?

Answer:

What amount (in mol) of Pb^{2+}(aq) ions remain in solution after the reaction?

Answer:

What is the final concentration of NO_3^-(aq) ions remaining in solution after the reaction?

Answer: