

Marks
5

- A solution is prepared by dissolving lead(II) nitrate (33.12 g) in 1.00 L of water. Write the balanced ionic equation for this dissolution reaction.

When a 100.0 mL portion of this solution is mixed with a solution of potassium iodide (0.300 M, 150.0 mL), a bright yellow precipitate of lead(II) iodide forms. Write the balanced ionic equation for this precipitation reaction.

What mass of lead(II) iodide is formed?

Answer:

What is the final concentration of Γ (aq) ions remaining in solution after the reaction is complete?

Answer: