constituent ions).

Marks • Depict the arrangement of water molecules around an ion. Explain why many ionic 3 compounds are soluble in water. negatively charged ion positively ? _____ charged ion Η .Ο、 [₩]_H^O_H `O´ н_о Η Θ Н Н ?+ ?+ Н dipolar water molecule positive end of water attracted negative end of water attracted to negatively charged ion to positively charged ion Water is a dipolar molecule. The positive ends (H) can directly interact with negative ions, while the negative end (O) can directly interact with positive ions. These interactions (called hydration enthalpy) are often sufficient to overcome the lattice enthalpy (the energy required to break up the ionic solid into its