Marks 7

- A 0.060 M solution of aluminium nitrate and a 0.080 M solution of potassium phosphate are prepared by dissolving $Al(NO_3)_3$ and K_3PO_4 in water. Write the ionic equations for these two dissolutions reactions.

Dissolution of $Al(NO_3)_3$	
Dissolution of K_3PO_4	

If these solutions are combined, aluminium phosphate precipitates. Write the ionic equation for the precipitation reaction.

100.0 mL of the aluminium nitrate solution is added to 50.0 mL of the potassium phosphate solution. What amount (in mol) of aluminium phosphate precipitates?

Answer:

What is the final concentration of aluminium ions remaining in solution after the precipitation?

Answer: