

	Marks
<ul style="list-style-type: none"><li>What is the pH of a 0.010 M solution of <math>\text{Ca}(\text{OH})_2</math>?</li></ul> <div data-bbox="705 555 1316 627" style="border: 1px solid black; padding: 5px; text-align: right;">pH =</div>	2
<ul style="list-style-type: none"><li>What is the pH of a 0.010 M solution of <math>\text{HNO}_2</math>? The <math>\text{p}K_a</math> of <math>\text{HNO}_2</math> is 3.15.</li></ul> <div data-bbox="705 1288 1316 1355" style="border: 1px solid black; padding: 5px; text-align: right;">pH =</div>	2
<ul style="list-style-type: none"><li>What is the pH of a solution that is 0.020 M in <math>\text{CH}_3\text{COOH}</math> and 0.010 M in <math>\text{CH}_3\text{CO}_2^-</math>? The <math>K_a</math> of <math>\text{CH}_3\text{COOH}</math> is <math>1.8 \times 10^{-5}</math> M.</li></ul> <div data-bbox="705 1948 1316 2016" style="border: 1px solid black; padding: 5px; text-align: right;">pH =</div>	2