

- Briefly explain why  $\text{H}_2\text{S}$  is a stronger Brønsted acid than  $\text{H}_2\text{O}$ .

**Marks**  
**2**

**S is much larger atom than O, so the H–S bond is much longer and weaker than H–O, so  $\text{H}_2\text{O}$  is weaker acid than  $\text{H}_2\text{S}$ .**