

Marks
8

- Solution A consists of a 0.50 M aqueous solution of HF at 25 °C. Calculate the pH of Solution A. The pK_a of HF is 3.17.

pH =

At 25 °C, 1.00 L of Solution B consists of 12.97 g of lithium fluoride, LiF, dissolved in water. Calculate the pH of Solution B.

pH =

Solution B (1.00 L) is poured into Solution A (1.00 L) and allowed to equilibrate at 25 °C. Calculate the pH of the final solution.

pH =

If you wanted to adjust the pH of the mixture of Solution A and Solution B to be exactly equal to 4.00, which component in the mixture would you need to increase in concentration?

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