

- You have completed a number of titrations during your laboratory work. What is the difference between the 'end point' and the 'equivalence point' in a titration?

**The equivalence point is the point where the reaction stoichiometry is exactly satisfied. For an acid base reaction, it corresponds to the point at which the amount of acid added is exactly equal to the amount of base initially present (or *vice versa*).**

**The endpoint is where the indicator changes colour and the reaction is observed to be completed.**

How do you need to consider that distinction when you chose an indicator for a particular titration?

**The endpoint and equivalence point need to be as close to one another as possible.**