

<ul style="list-style-type: none">Calculate the pH of a 0.020 M solution of $\text{Ba}(\text{OH})_2$.	Marks 1
<p style="text-align: right;">pH =</p>	
<ul style="list-style-type: none">Calculate the pH of a 0.150 M solution of HNO_2. The $\text{p}K_a$ of HNO_2 is 3.15.	3
<p style="text-align: right;">pH =</p>	
<ul style="list-style-type: none">Calculate the pH of a solution that is 0.080 M in acetic acid and 0.160 M in sodium acetate. The $\text{p}K_a$ of acetic acid is 4.76.	2
<p style="text-align: right;">pH =</p>	