

Calculate the pH of a 0.010 M solution of aspirin at 25 °C. The pK_a of aspirin is 3.5 at this temperature.

Marks
7

pH =

Aspirin, $C_9H_8O_4$ is not very soluble. “Soluble aspirin” can be made by reacting aspirin with sodium hydroxide. Write the chemical equation for this reaction.

Is a solution of “soluble aspirin” acidic or basic? Briefly explain your answer.