

- Calculate the pH of a 0.020 M solution of lactic acid, $\text{HC}_3\text{H}_5\text{O}_3$, at 25 °C. The $\text{p}K_a$ of lactic acid is 3.86.

pH =

A 1.0 L solution of 0.020 M lactic acid is added to 1.0 L of 0.020 M sodium hydroxide solution. Write the ionic equation for the reaction that occurs.

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Is the resulting solution acidic, basic or neutral? Give a reason for your answer.

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