
Marks
3

- Consider the values of the electronic energy levels of an He atom. State which interactions would be expected to increase the energies of the electrons and which would decrease them.

There are 3 electrostatic interactions:

- **Interaction between one electron and the 2 protons in the nucleus. This is attractive and lowers the energy of the electron.**
- **Interaction between the second electron and the 2 protons in the nucleus. This is attractive and lowers the energy of the electron.**
- **Interaction between the 2 electrons. This is repulsive and increases the energy of the electrons.**