Marks
6

- Magnesium hydroxide, $Mg(OH)_2$, is used as treatment for excess acidity in the stomach. Calculate the pH of a solution that is in equilibrium with $Mg(OH)_2$. The solubility product constant, K_{sp} of $Mg(OH)_2$ is 7.1×10^{-12} M^2.

ANSWER:

Determine whether 2.0 g of $Mg(OH)_2$ will dissolve in 1.0 L of a solution buffered to a pH of 7.00.

ANSWER: YES / NO