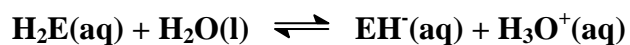


- Briefly explain why H₂Se is a stronger Brønsted-Lowry acid than H₂O and a weaker acid than H₂Te.

In general, when comparing binary acids within the same group, the strength of the bond E-H between the element (E) and hydrogen determines the acidity:



As the atomic size of E becomes larger, the E-H becomes longer and weaker. Thus H₃O⁺ is more readily formed in aqueous solution and acidity increases.