**Marks 4**

A solution is prepared that is 0.10 M in potassium bromide and 0.10 M in potassium chromate. A concentrated aqueous solution of silver nitrate is added with stirring. What is the concentration of $Ag^+(aq)$ ions when silver bromide first appears?
$K_{sp}$ of AgBr = $5.0 \times 10^{-13}$

Answer:

What is the concentration of $Ag^+(aq)$ ions when silver chromate first appears?
$K_{sp}$ of $Ag_2CrO_4$ = $2.6 \times 10^{-12}$

Answer:

What is the concentration of $Br^-(aq)$ ions when silver chromate first appears?

Answer: