• Draw the constitutional formulas of all isomers of C<sub>3</sub>H<sub>6</sub>BrCl.

Marks 6

A number of the above isomers are optically active. For all such compounds, draw the two enantiomers.

Cl 
$$H$$
  $Br$   $Cl$   $H$   $Br$   $Cl$   $(S)$ -1-bromo-1-chloropropane  $(S)$ -2-bromo-1-chloropropane  $(S)$ -2-bromo-1-chloropropane  $(S)$ -1-bromo-2-chloropropane  $(S)$ -1-bromo-2-chloropropane  $(S)$ -2-bromo-1-chloropropane  $(S)$ -2-bromo-1-chloropropane  $(S)$ -2-bromo-1-chloropropane

Select any **one** of the structures you have drawn on this page and write its full systematic name just below it. **See above.**