- Following blood donation, a solution of sodium oxalate is added to remove Ca^{2+} ions (as calcium oxalate, $CaC_2O_4 \cdot H_2O$, K_{sp} 2.3×10^{-9}), which cause the blood to clot. If the concentration of Ca^{2+} ions in blood is 9.7×10^{-5} g mL^{-1}, and 100.0 mL of 0.1550 M $Na_2C_2O_4$ is added to a 104 mL sample of blood, what will be the concentration (in mol L^{-1}) of Ca^{2+} ions remaining in the blood?

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Answer: