

Marks
6

- Solution A consists of a 0.050 M aqueous solution of benzoic acid, C_6H_5COOH , at 25 °C. Calculate the pH of Solution A. The pK_a of benzoic acid is 4.20.

pH =

Other than water, what are the major species present in solution A?

Solution B consists of a 0.050 M aqueous solution of ammonia, NH_3 , at 25 °C. Calculate the pH of Solution B. The pK_a of NH_4^+ is 9.24.

pH =

Other than water, what are the major species present in solution B?

THIS QUESTION CONTINUES ON THE NEXT PAGE.