

Marks
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- Complete the following table. Give, as required, the formula, the systematic name, the oxidation number of the underlined atom and, where indicated, the principal ions present in a solution prepared by adding the substance to water.

FORMULA	SYSTEMATIC NAME	OXIDATION NUMBER	PRINCIPAL IONS IN WATER SOLUTION
<u>N</u> O ₂			N/A
<u>Pb</u> (CH ₃ CO ₂) ₂			
			Mg ²⁺ (aq); <u>Cl</u> O ₄ ⁻ (aq)

Write the full electron configuration of the As³⁺ ion.

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- Draw the Lewis structures, showing all valence electrons for the following species. Indicate which of the species have contributing resonance structures.

HCO ₃ ⁻	COS	CN ⁻
Resonance: YES / NO	Resonance: YES / NO	Resonance: YES / NO