

- Consider the boiling points of the compounds 1-propanol, 1-propanethiol and 1-propaneselenol shown in the table below?

Compound	$\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$	$\text{CH}_3\text{CH}_2\text{CH}_2\text{SH}$	$\text{CH}_3\text{CH}_2\text{CH}_2\text{SeH}$
Boiling point ($^\circ\text{C}$)	97.2	67.8	147.0

With reference to intermolecular forces, explain briefly why the boiling points increase in the order $\text{CH}_3\text{CH}_2\text{CH}_2\text{SH} < \text{CH}_3\text{CH}_2\text{CH}_2\text{OH} < \text{CH}_3\text{CH}_2\text{CH}_2\text{SeH}$.