

**Marks**  
**5**

- Complete the following table. Give, as required, the formula, the systematic name, the oxidation number of the underlined atom and, where indicated, the number of *d* electrons for the element in this oxidation state.

Formula	Systematic name	Oxidation number	Number of <i>d</i> electrons
<u>C</u> O <sub>2</sub>	<b>carbon dioxide</b>	<b>+IV or +4</b>	<b>0</b>
Na <sub>2</sub> <u>Cr</u> O <sub>4</sub>	<b>sodium dichromate</b>	<b>+VI or +6</b>	<b>0</b>
<u>Fe</u> Cl <sub>3</sub> ·3H <sub>2</sub> O	<b>iron(III) chloride-3-water</b>	<b>+III or +3</b>	<b>5</b>
	potassium sulfate		