

- Complete the following table. Give, as required, the formula, the systematic name, the oxidation number of the underlined atom and, where indicated, the number of *d* electrons for the element in this oxidation state.

**Marks**  
**5**

Formula	Systematic name	Oxidation number	Number of <i>d</i> electrons
<u>C</u> O <sub>2</sub>			
Na <sub>2</sub> <u>Cr</u> O <sub>4</sub>			
<u>Fe</u> Cl <sub>3</sub> ·3H <sub>2</sub> O			
	potassium sulfate		