

Marks**7**

Calculate the partial pressure equilibrium constant, K_p , at 1200 K.

$K_p =$

What is the standard free energy change ΔG° for the forward reaction (in kJ mol^{-1}) at 1200 K?

$\Delta G^\circ =$

What will be the effect on the equilibrium if CO(g) is injected into the flask, which maintains a constant volume.

What will be the effect on the equilibrium if the temperature is decreased?

What will be the effect on the equilibrium if the volume of the flask is decreased?

What will be the effect on the equilibrium if the walls of the flask are refrigerated so that liquid water condenses out?