

<ul style="list-style-type: none"><li>Find the concentration of <math>\text{H}_3\text{O}^+</math> in a 0.60 M aqueous solution of nitrous acid. The acid dissociation constant of <math>\text{HNO}_2</math> is <math>K_a = 7.1 \times 10^{-4}</math> M.</li></ul>	<b>Marks</b> 2
Answer:	
<ul style="list-style-type: none"><li>An aqueous solution of a weak acid has <math>[\text{H}_3\text{O}^+] = 2.54 \times 10^{-4}</math> M. Find the pH and pOH of the solution.</li></ul>	1
pH =	pOH =
<ul style="list-style-type: none"><li>Ammonia, <math>\text{NH}_3</math>, is a Brønsted-Lowry base and a Lewis base, but not an Arrhenius base. Why?</li></ul>	3