

Marks
6

- Write the balanced half equations (including states) and the overall spontaneous reaction for a galvanic cell consisting of $\text{Ag}^+ | \text{Ag}$ and $\text{Sn}^{2+} | \text{Sn}$ half cells.

half equation at anode	
half equation at cathode	
overall reaction	

Express the overall reaction in voltaic cell notation.

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What is the sign of the cathode?

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What voltage would need to be applied to convert this galvanic cell into an electrolytic cell?

Answer: