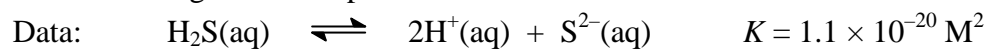


Marks
5

- The pH of a solution can be controlled by adding small amounts of gaseous HCl. Assuming no change in volume, calculate what the pH of the solution must be to just dissolve 1.00 g of NiS suspended in 1.0 L of water.



$$K_{\text{sp}}(\text{NiS}) = 1.0 \times 10^{-22} \text{ M}^2$$

pH =