

Marks
8

- An Ag electrode immersed in a saturated aqueous solution of AgBr has a reduction potential of 0.437 V at 25 °C with respect to the standard hydrogen electrode. Calculate the solubility product of AgBr at 25 °C.

Answer:

A Pd electrode immersed in an aqueous solution containing 0.01 Pd(NO₃)₂ M and 1.00 M NaCl has a reduction potential of -0.860 V at 25 °C with respect to the Ag electrode above. Calculate the stability constant of the complex ion, [PdCl₄]²⁻, at 25 °C.

Answer: