

<ul style="list-style-type: none">• Explain the meanings of the following terms.	Marks 8
<p>Heat</p> <p>Heat: energy contained in kinetic energies of molecules that flows from hotter to cooler temperatures.</p>	
<p>$P\Delta V$ work</p> <p>$P\Delta V$ work: work done by or on a system by a change in volume against a constant pressure.</p>	
<p>Internal energy</p> <p>Internal energy: the total energy contained within a system; the difference in internal energy in a system is the sum of the heat and work done by or on the system.</p>	
<p>Enthalpy change</p> <p>Enthalpy change: the difference in enthalpy between an initial and final state, the enthalpy being the heat of a system at constant pressure.</p>	
<p>Entropy</p> <p>Entropy: a measure of the distribution of heat, related to the number of ways or the probability of its distribution, hence to the level of disorder of the energy.</p>	
<p>Equilibrium constant</p> <p>Equilibrium constant: the ratio of the concentrations (or partial pressures) of reactants over products, each raised to its stoichiometric coefficient, when the system is at equilibrium.</p>	
<p>Reaction quotient</p> <p>Reaction quotient: the value of the equilibrium constant expression under any conditions, not at equilibrium.</p>	
<p>Triple point</p> <p>Triple point: The temperature and pressure at which a substance can exist as a solid, liquid and gas in equilibrium.</p>	