- Consider the following reaction.

$$
\mathrm{H}_{2} \mathrm{O}(\mathrm{~g})+\mathrm{Cl}_{2} \mathrm{O}(\mathrm{~g}) \rightleftharpoons 2 \mathrm{HOCl}(\mathrm{~g}) \quad K_{\mathrm{p}}=0.090 \text { at } 298 \mathrm{~K}
$$

Calculate $\Delta G^{\circ}$ (in $\mathrm{J} \mathrm{mol}^{-1}$ ) for this reaction.
$\square$
Calculate the reaction quotient, $Q$, at $25^{\circ} \mathrm{C}$ when $p\left(\mathrm{H}_{2} \mathrm{O}\right)=18 \mathrm{mmHg}$, $p\left(\mathrm{Cl}_{2} \mathrm{O}\right)=2.0 \mathrm{mmHg}$ and $p(\mathrm{HOCl})=0.10 \mathrm{mmHg}$.


In which direction will the reaction proceed spontaneously at these partial pressures?
$\square$

