

Marks
2

- A melt containing Cr^{3+} is electrolysed for exactly 1 hour with a current of 0.54 A. Calculate the quantity of chromium that is deposited in this time at the electrode.

Answer:

4

- An Ag electrode immersed in an aqueous solution containing AgNO_3 (0.010 M) and NaCN (1.00 M) has a potential of -0.66 V. Calculate the stability constant of the complex ion, $[\text{Ag}(\text{CN})_2]^-$.

Answer: