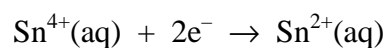
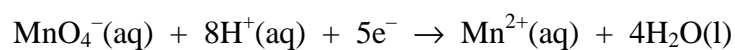


- Give the oxidation number of carbon in each of the following.

$\text{CF}_2\text{Cl}_2(\text{g})$	
$\text{Na}_2\text{C}_2\text{O}_4(\text{s})$	
$\text{HCO}_3^-(\text{aq})$	
$\text{C}(\text{s})$	

**Marks**  
**2**

- Consider a voltaic cell that uses the following half-reactions:



Write a balanced equation for the overall reaction.

**3**

Which species is the oxidising agent?

Which species is the reducing agent?

Calculate the standard cell potential. (Refer to the table of standard reduction potentials.)

Answer: