CHEM1405 2008-J-9 June 2008 22/25(a)

• A buffer solution is formed with 0.250 M CH ₃ COOH and 0.350 M CH ₃ COONa. What is the pH of this buffer solution? (K_a of acetic acid = 1.8×10^{-5} M.) Answer: Calculate the pH of the solution formed when 6.3×10^{-2} mol of NaOH is added to 1.0 L of the buffer solution.	•	Codeine, a cough suppressant extracted from crude opium, is a weak base with a $pK_b = 5.79$. What is the pH of a 0.020 M solution of codeine?		
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