

Marks
6

- Write the balanced chemical equation for the dissolution of solid $\text{Ca}(\text{CH}_3\text{CO}_2)_2$ in water.

What is the pH of a solution that has 158.2 g of $\text{Ca}(\text{CH}_3\text{CO}_2)_2$ dissolved in 1.000 L of water? The $\text{p}K_a$ of acetic acid, CH_3COOH , is 4.76.

pH =

Calculate the pH of this solution after the addition of 0.250 mol of HCl gas?

pH =