

- Both $\text{HCO}_3^-(\text{aq})$ and $\text{CO}_2(\text{aq})$ are present in human blood. How does their presence ensure that the pH of blood is maintained at ~ 7.2 , even if $\text{H}^+(\text{aq})$ or $\text{OH}^-(\text{aq})$ are produced by processes in the body?

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How does hyperventilation (very rapid breathing) interfere with this balance? What is the effect?