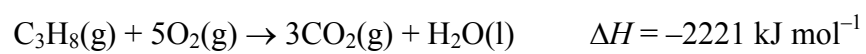
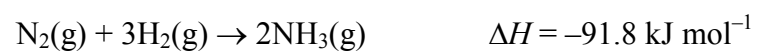


- What quantity of heat is released when 15.2 g of propane (C_3H_8) is burnt according to the following equation?

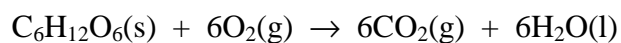
**3**

- How much heat is evolved when 907 g of ammonia is produced according to the following equation? (Assume the reaction occurs at constant pressure.)



Answer:

- The balanced equation for the complete oxidation of glucose to carbon dioxide and water is given below.



Calculate the mass of carbon dioxide produced by the complete oxidation of 1.00 g of glucose.

Marks
3

Answer:

Calculate the volume of this mass of carbon dioxide at 0.50 atm pressure and 37 °C.

Answer: