

<ul style="list-style-type: none">In the spaces provided, briefly explain the meaning of the following terms.	Marks 3
<p>Effective nuclear charge</p> <p>The force of attraction experienced by the outer electrons of an atom. It's a combination of the magnitude of the nuclear charge mitigated by shielding by the inner electrons. Effective nuclear charge increases to the top right of the periodic table.</p>	
<p>Atomic emission spectrum</p> <p>Unique for each element, the spectrum represents emission of light of discrete frequencies corresponding to the energy differences between electron energy levels in an atom.. It results from the movement of electrons from a higher energy level to a lower one.</p>	
<p>Core electrons</p> <p>Inner shell electrons that are not involved in bonding.</p>	

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