

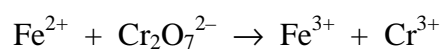
Marks
2

- A concentration cell containing aqueous solutions of $\text{Cu}(\text{NO}_3)_2$ and solid copper metal is constructed so that the Cu^{2+} ion concentration in the cathode half-cell is 0.66 M. Calculate the concentration of the Cu^{2+} ion in the anode half-cell if the cell potential for the concentration cell at 25 °C is 0.03 V.

Answer:

3

- In **acid solution**, dichromate ion oxidises iron(II) to iron(III) as illustrated in the partial equation:



Write a balanced equation for this reaction.

What would happen to the cell potential if the concentration of Cr^{3+} were increased?