- Give the equilibrium concentration of Ni^{2+}(aq) ions in a solution formed by dissolving 0.15 mol of $NiCl_2$ in 0.500 L of 2.00 M KCN solution. The K_{stab} of $[Ni(CN)_4]^{2-} = 1.7 \times 10^{30}$.

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Answer: