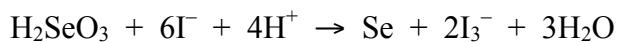


- The following reaction is run from 4 different starting positions.



Marks
4

Experiment	Initial $[\text{H}_2\text{SeO}_3]$ (mol L ⁻¹)	Initial $[\text{I}^-]$ (mol L ⁻¹)	Initial $[\text{H}^+]$ (mol L ⁻¹)	Initial rate of increase of $[\text{I}_3^-]$ (mol L ⁻¹ s ⁻¹)
1	0.100	0.100	0.100	1.000
2	0.100	0.075	0.100	0.422
3	0.075	0.100	0.100	0.750
4	0.100	0.075	0.075	0.237

Determine the rate law for the reaction.

What is the value of the rate constant?

Answer: