

- The  $K_{sp}$  of AgBr is  $5.0 \times 10^{-13}$ . The  $K_{stab}$  of  $[\text{Ag}(\text{S}_2\text{O}_3)_2]^{3-}$  is  $4.7 \times 10^{13}$ . Calculate the value of the equilibrium constant for the dissolution of AgBr in  $\text{Na}_2\text{S}_2\text{O}_3$  solution.

**Marks**  
**7**

Answer:

Calculate the solubility of AgBr in 2.0 M  $\text{Na}_2\text{S}_2\text{O}_3$ .

Answer:

The  $K_{stab}$  for  $[\text{Ag}(\text{S}_2\text{O}_3)_2]^{3-}$  is much greater than the  $K_{stab}$  for  $[\text{Ag}(\text{NH}_3)_2]^+$ . Explain why this is so.