Marks
4

- At a certain temperature the following data were collected for the reaction shown.

$$2ICl + H_2 \rightarrow I_2 + 2HCl$$
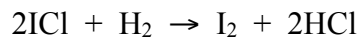
Experiment	Initial [ICl] (mol L^{-1})	Initial [H$_2$] (mol L^{-1})	Rate of formation of [I$_2$] (mol L^{-1} s^{-1})
1	0.10	0.10	0.0015
2	0.20	0.10	0.0030
3	0.10	0.050	0.00075

Determine the rate law for the reaction.

What is the value of the rate constant?

Answer:

THE REMAINDER OF THIS PAGE IS FOR ROUGH WORKING ONLY.