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- The freezing point of a sample of seawater is measured as $-2.15\text{ }^{\circ}\text{C}$ at 1 atm pressure. Assuming that the concentrations of other solutes are negligible, determine the molality (in mol kg^{-1}) of NaCl in this sample. The molal freezing point depression constant for H_2O is $1.86\text{ }^{\circ}\text{C kg mol}^{-1}$.

Marks
3

Answer:

THE REMAINDER OF THIS PAGE IS FOR ROUGH WORKING ONLY.