Marks

• Consider the following reaction:		3
$2N_2O(g) + 3O_2(g) \rightarrow 4NO_2(g)$		
Calculate ΔG° for this reaction given the following data.		
$4NO(g) \rightarrow 2N_2O(g) + O_2(g)$	$\Delta G^{\circ} = -139.56 \text{ kJ mol}^{-1}$	
$2NO(g) + O_2(g) \rightarrow 2NO_2(g)$	$\Delta G^{\circ} = -69.70 \text{ kJ mol}^{-1}$	
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	Answer:	
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THE REMAINDER OF THIS PAGE IS FOR ROUGH WORKING ONLY.