	Marks
• The solubility product constant of $Fe(OH)_3$ is 1×10^{-39} M^4. What is the concentration of $Fe^{3+}(aq)$ in equilibrium with $Fe(OH)_3$ at pH 7.0?	4

ANSWER:

To what value does the pH need to be increased to decrease the concentration of $Fe^{3+}(aq)$ to a single $Fe^{3+}(aq)$ ion per litre of solution?

ANSWER: