	Marks
• Calcium oxalate is a major constituent of kidney stones. Calculate the solubility product constant for calcium oxalate given that a saturated solution of the salt can be made by dissolving 0.0061 g of $CaC_2O_4 \cdot H_2O(s)$ in 1.0 L of water.	**2**
Answer:	
• A sample of 2.0 mg of $Cu(OH)_2$ is added to 1.0 L of a solution buffered at a pH of 8.00. Will all of the $Cu(OH)_2$ dissolve? Show all working. (The K_{sp} of $Cu(OH)_2$ is 4.8×10^{-20} M^3.)	**3**
Answer:	