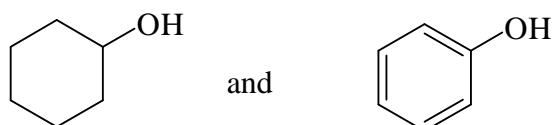
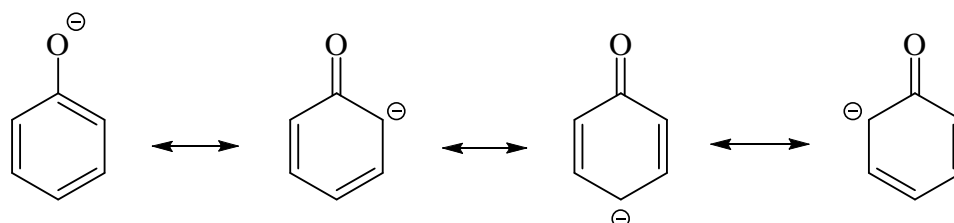


Marks
4

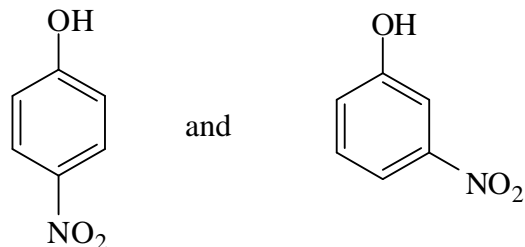
- For each of the following pairs of compounds, identify which is the stronger acid and give reasons for your choice.



Phenol is more acidic as the phenoxide ion is resonance stabilised:



Chloroacetic acid is more acidic. The Cl atom is electronegative and pulls electrons from the carboxylic acid, thus weakening the O–H bond.



The *para* isomer is more acidic as the charge on the phenoxide ion can be delocalised into the nitro group. This is not possible with the *meta* isomer.

