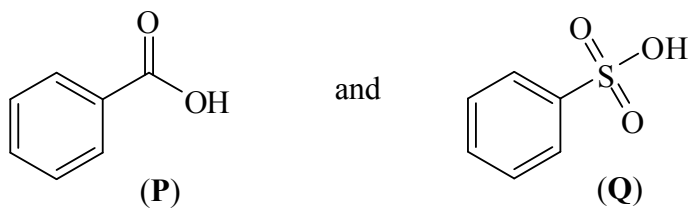
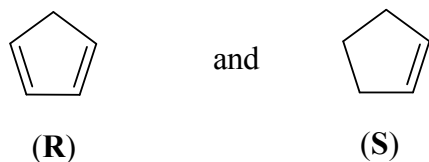
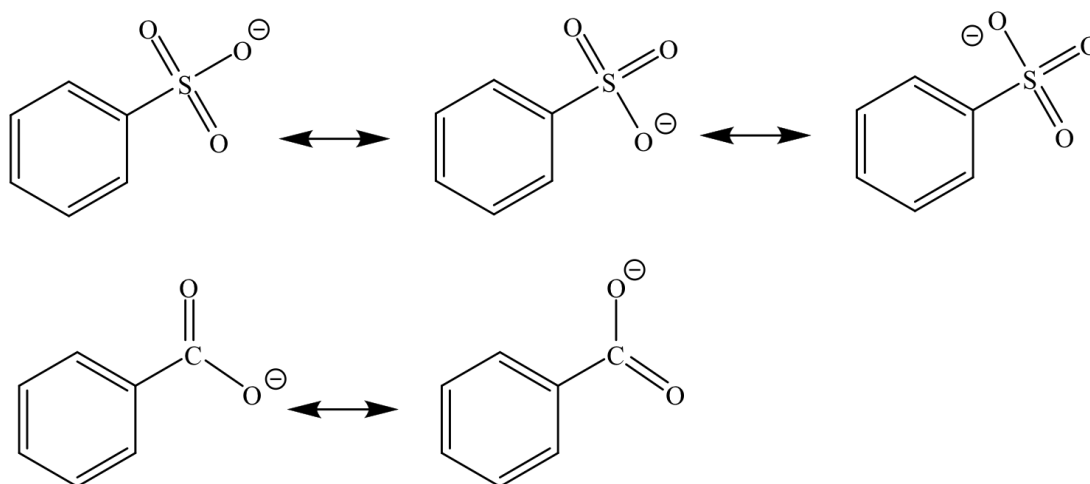


- For each of the following pairs of compounds, identify which is the stronger acid and give reasons for your choice.

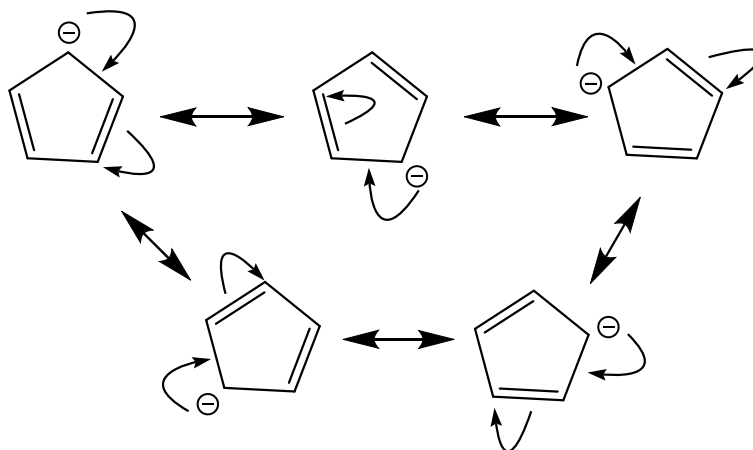
Marks
3



(Q) There is greater resonance stabilisation of the conjugate base (more canonical forms):



(R) There is greater resonance stabilisation of the conjugate base because it is aromatic.



ANSWER CONTINUES ON THE NEXT PAGE

CF₃CO₂H and CH₃CO₂H

(T)

(U)

(T) There is greater resonance stabilisation of the conjugate base due to the inductive electron withdrawal of the very electronegative F atoms.

