• Aqua ligands in coordination complexes are generally acidic. Briefly explain this phenomenon using $\left[\text{Co(NH}_3)_5(\text{OH}_2)\right]^{3+}$ as an example.		Ma
Solution A consists of a 0.10 M aqueous at 25 °C. Calculate the pH of Solution A	solution of $[Co(NH_3)_5(OH_2)](NO_3)_3$ A. The p K_a of $[Co(NH_3)_5(OH_2)]^{3+} = 5.69$.	
	pH =	
At 25 °C, 1.00 L of Solution B consists of in water. Calculate the pH of Solution B	of 28.5 g of $[Co(NH_3)_5(OH)](NO_3)_2$ dissolved 3.	
	pH =	
Using both Solutions A and B, calculate a $1.0 L$ solution with a pH = 7.00 .	the volumes (in mL) required to prepare	